

Mr. Lin

Name \_\_\_\_\_

Per \_\_\_\_\_

## Solubility and Naming Compounds

AP Chemistry

AP Chemistry

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### Name the following compounds from their formula:.

1.  $\text{NH}_4\text{I}$
2.  $\text{Fe}_2(\text{SO}_3)_3$
3.  $\text{H}_3\text{PO}_4$  (aq)
4.  $\text{AgClO}_3$
5.  $\text{SnC}_2\text{O}_4$
6.  $\text{NaCN}$
7.  $\text{Zn}(\text{OH})_2$
8.  $\text{Rb}_2\text{SiO}_3$
9.  $\text{MnO}_2$
10.  $\text{H}_2\text{O}_2$
11.  $\text{KCl}$
12.  $\text{KC}_2\text{H}_3\text{O}_2$
13.  $\text{CaSO}_4$
14.  $\text{Sb}_2(\text{Cr}_2\text{O}_7)_3$
15.  $\text{Li}_3\text{P}$

### Write the chemical formula from the compounds name:

1. Lead (IV) Carbonate
2. Magnesium Fluoride
3. Cesium Arsenate
4. Boron (III) Silicate
5. Francium Nitrite
6. Cobalt (II) Chloride
7. Mercury (I) Permanganate

8. Mercury (I) Fluoride
9. Barium Phosphate
10. Bismuth (III) Hydride
11. Aluminum Borate
12. Gold Oxide
13. Copper (I) Hypochlorite
14. Tin (II) Chlorite
15. Phosphorous Pentachloride

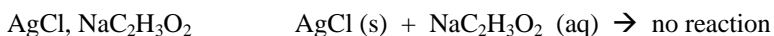
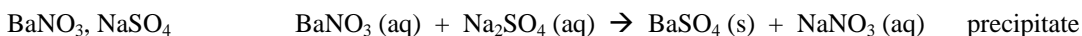
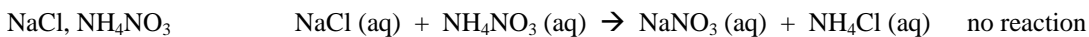
## Solubility

**Directions: Two beakers are filled with 250 ml of water. Determine if the following compounds will dissolve in each beaker and if a precipitate will form if both beakers are mixed.**

(aq) = aqueous → soluble

(s) = solid → insoluble

Examples:



### **Beaker 1/ Beaker 2**

- 1) NaCl, Ba(NO<sub>3</sub>)<sub>2</sub>
- 2) SrSO<sub>4</sub>, NH<sub>4</sub>ClO<sub>4</sub>
- 3) NaCl, AgNO<sub>3</sub>
- 4) Au<sub>2</sub>CO<sub>3</sub>, PbSO<sub>4</sub>
- 5) Ag<sub>2</sub>S, HCH<sub>3</sub>CO<sub>2</sub>
- 6) Cs<sub>2</sub>SO<sub>4</sub>, SrCl<sub>2</sub>
- 7) Fe<sub>2</sub>O<sub>3</sub>, (NH<sub>4</sub>)<sub>3</sub>P
- 8) CaF<sub>2</sub>, MgBr<sub>2</sub>
- 9) Ag<sub>2</sub>SO<sub>4</sub>, PbI<sub>2</sub>
- 10) (NH<sub>4</sub>)<sub>2</sub>CO<sub>3</sub>, NaI