

Names, Formulas, and Charges of Common Ions

Positive Ions (Cations)

1+	2+	3+	4+
Ammonium NH_4^+ Cesium Cs^+ Gold Au^+ Silver Ag^+	Nickel(II) Ni^{2+} Zinc Zn^{2+}	Antimony(III) Sb^{3+} Arsenic(III) As^{3+} Bismuth(III) Bi^{3+} Boron(III) B^{3+} Titanium Ti^{3+}	
<u>Copper(I) $\text{Cu}^+ \rightarrow$</u>	<u>Copper(II) Cu^{2+}</u> <u>Cobalt(II) $\text{Co}^{2+} \rightarrow$</u> <u>Chromium(II) $\text{Cr}^{2+} \rightarrow$</u> <u>Iron(II) $\text{Fe}^{2+} \rightarrow$</u> <u>Lead(II) $\text{Pb}^{2+} \rightarrow$</u> <u>Manganese(II) $\text{Mn}^{2+} \rightarrow$</u>	<u>Cobalt(III) Co^{3+}</u> <u>Chromium(III) Cr^{3+}</u> <u>Iron(III) Fe^{3+}</u>	<u>Lead(IV) Pb^{4+}</u> <u>Manganese(IV) Mn^{4+}</u>
<u>Mercury(I) $\text{Hg}_2^{2+} \rightarrow$</u>	<u>Mercury(II) Hg^{2+}</u> <u>Tin(II) $\text{Sn}^{2+} \rightarrow$</u>		<u>Tin(IV) Sn^{4+}</u>

Negative Ions (Anions)

1-	2-	3-
Acetate $\text{C}_2\text{H}_3\text{O}_2^-$ Hypochlorite ClO^- Chlorite ClO_2^- Chlorate ClO_3^- Perchlorate ClO_4^- Nitrite NO_2^- Nitrate NO_3^- Carbide C_2^{2-} Hydride H^- Cyanide CN^- Hydroxide OH^- Thiocyanate SCN^- Permanganate MnO_4^-	Silicate SiO_3^{2-} Peroxide O_2^{2-} Chromate CrO_4^{2-} Dichromate $\text{Cr}_2\text{O}_7^{2-}$	Arsenate AsO_4^{3-} Borate BO_3^{3-} Nitride N^{3-}
<u>Dihydrogen phosphate $\text{H}_2\text{PO}_4^- \rightarrow$</u> <u>Hydrogen oxalate $\text{HC}_2\text{O}_4^- \rightarrow$</u> <u>Hydrogen carbonate $\text{HCO}_3^- \rightarrow$</u> <u>Hydrogen sulfide $\text{HS}^- \rightarrow$</u> <u>Hydrogen sulfite $\text{HSO}_3^- \rightarrow$</u> (Bisulfite) <u>Hydrogen sulfate $\text{HSO}_4^- \rightarrow$</u> (Bisulfate)	<u>Hydrogen phosphate $\text{HPO}_4^{2-} \rightarrow$</u> <u>Oxalate $\text{C}_2\text{O}_4^{2-}$</u> <u>Carbonate CO_3^{2-}</u> <u>Sulfide S^{2-}</u> <u>Sulfite SO_3^{2-}</u> <u>Sulfate SO_4^{2-}</u> Thiosulfate $\text{S}_2\text{O}_3^{2-}$	Phosphide P^{3-} Phosphite PO_3^{3-} <u>Phosphate PO_4^{3-}</u>